

## **Mercury Removal from Aqueous Solution by Mixed Mineral Systems I. Reactivity and Removal Kinetics**

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**Abstract :** *This study investigates the reactivity and removal kinetics of arsenite on mixed mineral systems from aqueous solution related to contaminated waters. The sorbents used were kaolinite, montmorillonite, goethite, and their mixtures. The effects of, proton coefficient, and sorption kinetics were studied at room temperature (23 ± 2 °C). Reactivity studies demonstrate enhancement of proton coefficient and the acidity of reactive sites by mixed mineral systems except for kaolinite-montmorillonite, thus increasing mercury removal by proton exchange. Kinetic studies demonstrate two- phase reactions attributed to outer sphere complexation and inner sphere complexation. In the first-phase reaction, mineral mixing based on our empirical model decrease the mass transfer rates for kaolinite-montmorillonite and montmorillonite-goethite, not affecting kaolinite-goethite. For the second-phase reaction, mineral mixing did not change the mass transfer rates for the mixed mineral systems except for montmorillonite-goethite. The behaviors of the mixed mineral systems in mercury sorption suggest that different reactive sites were involved at the onset of sorption, with reactions and sorption ending with inner-sphere complexation. Differences between actual and theoretical % sorption progressed from negative to positive for all mixed minerals as contact time was increased. This means that increase in contact time during Hg-sorbent interaction could help increase mercury removal based on mineral mixing.*

**Keywords:** *mercury, removal, reaction, kinetics, mixed mineral systems.*

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### **I. Introduction**

Mercury poses an environmental and human threat due to its toxicity at low concentrations. It occurs in three forms namely (Hg<sup>0</sup>), (Hg<sup>2+</sup>) and (HgP) and possesses high toxicity, volatility. Mercury bioaccumulates in the environment and creates neurological health impact [1]. Moreover, it appeared as a critical and chronic problem because it can easily be transformed into methyl mercury, an organic form, by bacteria in bottom level sediments which is taken up by organisms more rapidly and is much more toxic and stable than inorganic form [2].

Mercury and its compounds recognized as dangerous and insidious poisons and can be adsorbed through the gastrointestinal tract and also through the skin and lungs [3] which is stored in the liver, kidneys, brain, spleen and bone leading to development of carcinogenic, mutagenic, teratogenic and also promotes several health problems, tyrosinemia, paralysis, serious intestinal and urinary complications, dysfunction of the central nervous system and in more severe cases of intoxication, death. The severity of mercury toxicity was recognized in the late 1950s and 1960s because of an environmental tragedy in Minamata Bay, Japan, where hundreds of individuals suffered from mercury poisoning (Minamata disease) [4-5]

As a common heavy metal, mercury has been used in medical treatment and industry, but its toxic properties for human health and ecosystem have attracted increasing attention [6-8]. Sorption kinetic experiments revealed that mercury sorption is a relatively rapid process, where film diffusion and intra-particle diffusion play an important role. The US Environmental Protection Agency (USEPA) prepared a list of 129 organic and inorganic pollutants found in wastewater that constitute serious health hazards [9] This Priority Pollutants List includes thirteen elements: Sb, As, Be, Cd, Cr, Cu, Pb, Hg, Ni, Se, Ag, Tl and Zn [10-12].

For successful treatment of water containing mercury using mineral systems, a full understanding of the reactivity and the kinetic reaction partners to the reactive sites of mercury sorbents is required [13-14]. The removal of mercury from aquatic systems is controlled by the speciation, mobility, ionic size of the sorbing ions, sorbate composition, sorbent solubility, sorbent particle size, sorbent surface charge, surface area of the mineral sorbent, solution dilution and H<sup>+</sup>/ – M exchange stoichiometry [15-16].

The sorption of a solute present in a solid-solution system is usually assumed to consist of four consecutive steps: external mass transfer of sorbate molecules across the boundary liquid film, binding of sorbate molecules on the active sites on the surface of the sorbent, intra-particle diffusion of sorbate molecules, and sorption of sorbate molecules on the active sites distributed within the sorbent particles. Reactivity of clay particles are associated with OM, Fe or Al oxides [17-18] and combined with elevated temperatures and acidic pH (4–5) favor geochemical conditions and nutrient stocks favorable for Hg bacterial methylation [19]. The

rapid detection and removal of  $Hg^{2+}$  from pollutants are dependent on a full understanding of the reactivity and removal kinetics of mercury in aqueous solution [20-25].

Several methods have been established in the treatment of mercury contaminated water. Some of these techniques include bamboo charcoal, apatite IITM, iron-based sorbents, diperiodatocuprate(III) coordination solution, malt spent rootlets, bacillus cereus, modified  $Fe_3O_4$ , flocculation method, complexation by dissolved organic matter, heterogeneous photocatalysis with  $TiO_2$ , alginate and pectate [[26-35]. Others have reviewed the technologies for removing mercury in aqueous solution [36]. However, some of these new techniques are rather expensive for limited size water treatment systems situated in rural communities and the reactivity and removal kinetics require further investigation. Consequently, innovative cost-effective treatment processes are urgently needed. One of such emerging method is the use of mixed mineral systems of clays and hydroxide(s). Mixed mineral systems that could be applied for mercury removal from contaminated water are readily available locally.

### 1.1.Theoretical models and isotherms

To addresses the reactivity and removal kinetics involved in sorbent-sorbate interaction, a theory is designed to explain and predict the behavior of mineral-Hg interactions. Adsorption isotherms give a quantitative relationship between the solute concentration in the solution and the amount of solute adsorbed per unit mass of the adsorbent surface at equilibrium conditions. Langmuir, Freundlich, Dubinin–Radushkevich (D–R) are the commonly used isotherms [37]. In the present study, an empirical model derived from Freundlich isotherm model was chosen.

Detailed system characterization and an empirical model involving the distribution coefficient (Kd) as used in this paper are provided elsewhere [38].  $Kd(L/kg)$  was calculated from the Freundlich model equation,

$$S = KdC^N \tag{1}$$

where S is the sorbed concentration ( $\mu g/kg$ ), Kd is the distribution coefficient, C is the equilibrium concentration ( $\mu g/l$ ), and  $N = 1$  is a chemical-specific coefficient derived from the slope of the plot. The empirical model to address the mineral-Hg interactions is given:

$$Kd_{total} = Kd_{total} = \frac{Kd_1 + Kd_2 + Kd_n}{n} \tag{2}$$

where  $Kd_{total}$  is the theoretical distribution coefficient for a 1:1 mixed suspension,  $Kd_1$  is the distribution coefficient for first single mineral suspension, and  $Kd_2$  is the distribution coefficient for second single mineral suspension,  $Kd_n$  is the distribution coefficient for n number of mineral suspensions and n is the number of mineral suspensions.

The simple empirical model used for the partitioning of a sorbed mercury contaminant between single mineral phases and mixed mineral phases is based on the assumptions that the following could account for differences between single and mixed mineral sorption:

1. Secondary mineral phase developed during sorbate-sorbent interaction .
2. Components of minerals in the mixed mineral suspension acted as chemisorbed species and not as individual networks. 3. differential mass of mixed and single mineral phases.

The difference between the actual sorption and the theoretical sorption was used to clarify the effects of mineral mixing on Hg(II) sorption. Mineral mixing is said to (a) enhance Hg(II) removal where the difference is positive; (b) depresses or attenuate Hg(II) removal where the difference is negative; and (c) have no effect on Hg(II) removal where no difference exist between Hg(II) sorbed and theoretical Hg(II) sorption.

The difference between the actual Kd and the theoretical ( $Kd_{total}$ ) was used to clarify the effects of mineral mixing on Hg removal. Mineral mixing is said to (a) enhances Hg removal where the difference is positive; (b) attenuate Hg removal where the difference is negative; and (c) have no effect on Hg removal where no difference exist between the actual Kd and theoretical Kd:

$$\Delta Kd = Kd - Kd_{total} \tag{3}$$

For the reactivity and removal kinetic studies, the empirical model for the mixed mineral systems was related to  $\alpha$  and  $K_f$  as follows,

$$\Delta \alpha = \alpha - \alpha_{total} \tag{4}$$

$$\Delta K_f = K_f - K_{f total} \tag{5}$$

where  $\alpha$ ,  $K_f$ ,  $\alpha_{total}$  and  $K_{ftotal}$ <sup>1</sup> are the proton coefficient, mass transfer rate,

theoretical proton coefficient and theoretical mass transfer coefficient, respectively.

The main objective of this work is to determine the effects of mineral mixing on the reactivity and kinetics involved in mercury removal from contaminated water. The sorption kinetic model assumed that sorption rate was determined by sorption interaction between the sorbent reactive sites and the sorbate involving outer sphere complexation and inner sphere complexation [39]. Otherwise, the intra-particle diffusion involving the diffusion of the adsorbate in the pore of the adsorbent as a third sorption reaction step was involved [40]. This is due to the fact that surfaces of clays and hydroxides have a high concentration of OH<sup>-</sup> groups readily protonated at low pH, generating positive charge which retains anions by electrostatic attraction [41].

The mass balance of Hg(II) sorbed per unit mass of the mixed mineral suspension (mg/g) was calculated by the following [42-44]:

$$Q_e = \frac{(C_i - C_e)V}{W} \tag{6}$$

where  $C_i$  and  $C_e$  are the initial and equilibrium Hg(II) concentrations in mg/l.,  $V$  is volume of the Hg(II) solution in ml, and  $W$  is the weight of adsorbent in mg respectively.

## II. Materials and Methods

### 2.1. System characterization

Clays used in this study provided by the Richard Baker Harrison Company and Acros Organics Ltd and (hydr)oxides provided by Iconofile Company Inc were nitrogen flushed and stored in airtight containers to avoid surface oxidation. Mercuric chloride (HgCl<sub>2</sub>) provided by Iconofile Company Inc was employed as the source of Hg(II). A standard solution containing variable concentrations of Hg(II) in ppm was prepared by dissolving HgCl<sub>2</sub> (Merck) in distilled water. In all experimental studies conducted in triplicates, samples were stored in the dark at room temperature (23±3 °C) not exceeding 24 h before analysis [45]. Supernatant was filtered through a cellulose acetate filter (pore size 0.2µm) and analyzed for Hg(II) using a Hitachi Atomic Absorption Spectrophotometer (HG-AAS). The working solutions of different concentrations were prepared by diluting the stock solution immediately before starting the batch studies.

For sorbent characterization, the (a) Coulter laser method was used to determine the particle sizes; (b) % colloid was estimated from the particle size distribution curves; (c) equilibrium pH of the untreated mineral suspensions was determined using the Model 3340 Jenway ion meter; (d) the standard volumetric Brunauer, Emmett, and Teller (BET) method was used to determine the surface areas [46-47] (e) the Na saturation method was used to determine cation exchange capacity (CEC) of the clay minerals[5] (Table 1). (f) spectral analysis was performed using scanning electron microscopy (SEM) and energy dispersive spectroscopy (EDS) to confirm the mineral sorbent[48-49].

### 2.2. Reactivity experiments

For reactivity studies to determine the proton coefficient as provided (Eqs. (8), (9), standard laboratory procedure was used [50-51]. 1% single and 1:1 mixed mineral suspensions with no added electrolyte were reacted with solution containing 10ppm of arsenite regulated to the required pH at the start of experiments. To validate the sorption mechanism involved in mercury removal, 1% single and 1:1 mixed mineral suspensions made up to 50 ml were reacted with solution containing 10 ppm of Hg(II) regulated to pH 4. Supernatant was filtered through a cellulose acetate filter (pore size 0.2 µm) and analyzed for Hg(II), using a Hitachi Atomic Absorption Spectrophotometer (HG-AAS).

**Table1: Characteristics of clays and hydrous oxides [49]**

Mineral	Particle size (µm)	% (<1 µm) colloid	pH ± σ	Surface area(SSA±σ) (m <sup>2</sup> /g)
Kaolinite	20.01±0.5	3.00	6.05±0.05	47.01± 0.24
Montmorillonite	80.05±0.20	0.53	2.01±0.09	10.00± 0.00
Goethite	40.10±0.15	2.92	8.05±0.06	71.05± 0.17
Kaolinite/montmorillonite	80.05±50	0.97	5.01±0.02	88.05± 0.55
Montmorillonite/goethite	15.25±0.24	3.85	3.03± 0.04	147.10± 0.50
Kaolinite/goethite	140.35±55	0.73	3.05± 0.01	79.30± 0.59

Surface charge controls the reactivity of mineral surfaces and is dependent on the acid–base properties of surface hydroxyl groups, where H<sup>+</sup> and OH<sup>-</sup> are the potential-determining ions [26]. Sorption of Hg(II) on mineral surfaces requires proton exchange, the stoichiometry of this reaction is described [49] and the proton consumption function is given by:



$$\log K_d = \log K_p \{ \text{SOH} \}^{\alpha + \alpha \text{pH}}, \tag{8}$$

where SOH is the mineral surface-binding site, M<sup>2+</sup> is the soluble metal species, (SOH)<sub>α</sub> – M is the surface bound metal, logK<sub>p</sub> is the apparent equilibrium binding constant, and α is the proton coefficient, representing the number of protons displaced when one mole of metal binds to the mineral surface [[32]. Proton coefficient was calculated from the slope of logK<sub>d</sub> versus pH plot (Fig. 2),

### Kinetic experiments

For Hg(II) removal kinetics experiments, 1% single and 1:1 mixed mineral systems were reacted with solution containing 10ppm of Hg(II) regulated to pH 4. Amounts of Hg(II) remaining in solution after 2, 4, 6, 8, 12, 18, and 24 h were determined using Hitachi Atomic Absorption Spectrophotometer (HG-AAS). Twenty-four hours was sufficient for kinetic studies because sorption reactions occur in milliseconds or minutes [45, 48]. The transport of adsorbate from external layers to the mineral surface where sorption occurs is dependent on a mass transfer constant K<sub>f</sub> obtained from the slopes of the curve derived from plotting C<sub>t</sub>/C<sub>0</sub> vs time [34, 49],

$$\left[ \frac{d(C_t/C_0)}{dt} \right]_{t=0} \cong -K_f S_s \tag{9}$$

where C<sub>t</sub> and C<sub>0</sub> are the initial concentrations of Hg(II) at time t, S<sub>s</sub> is the exposed external surface area of the sorbent, and K<sub>f</sub> is the mass transfer coefficient [42]. The Freundlich isotherm was chosen to describe sorption of Hg(II) because this is suitable for heterogeneous surfaces over a wide range of solute concentrations [34, 43].

## III. Results And Discussion

### 3.1. Mixed mineral systems and H<sup>+</sup>/-M exchange stoichiometry

Although the proton coefficient (α) is not characteristic of a particular mineral (Table 2, derived from Figs.1), differences in α for mixed mineral systems compared to each other and compared to single mineral systems may be linked to differences in the availability of strongly acidic sites All proton coefficient for Hg(II) sorbed on single mineral systems except for goethite are greater than one. This indicated high level of protonation during the sorption process. Proton coefficient for Hg(II)-kaolinite interaction was higher than Hg(II) sorbed on goethite and montmorillonite. This may be attributed to the acidic sites present on kaolinite planar surfaces. Significant numbers of weakly acidic edge sites in goethite and montmorillonite may reduce the proton coefficient because of the limited exchange of protons for sorbing ions.

**Table2: Proton coefficients (α) and regression coefficient (R<sup>2</sup>) of Hg(II) sorbed on mineral suspensions**

Mineral suspensions	R <sup>2</sup>	α	α <sub>total</sub>
Kaolinite	0.99	1.32	⊗
Goethite	0.99	0.69	⊗
Montmorillonite	0.99	1.25	⊗
Kaolinite/montmorillonite	0.99	0.77	1.28
Goethite/kaolinite	0.99	1.24	1.00
Goethite/montmorillonite	0.99	1.33	0.97

Note: ⊗ not applicable

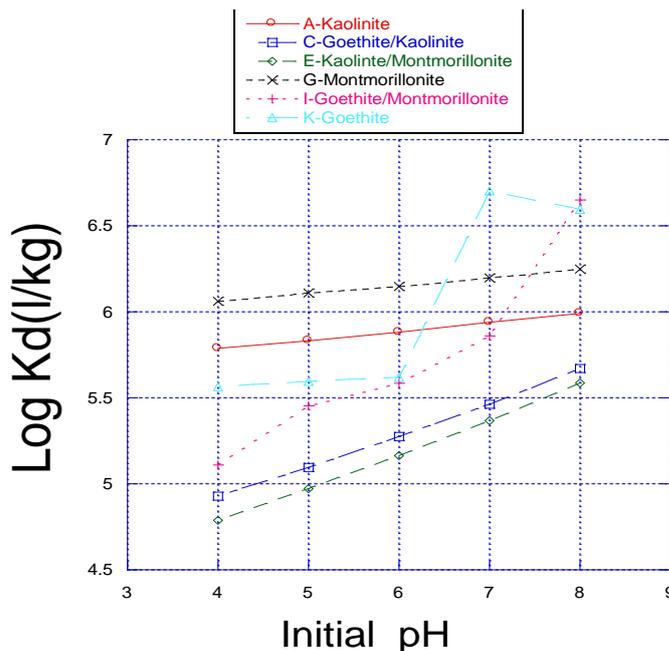


Figure 1: Plots of logKd versus Initial pH for Hg(II) sorbed on single and mixed mineral systems.

Except for kaolinite/montmorillonite, mixed mineral suspension,  $\alpha_{total}$  for Hg(II) sorbed on the mixed mineral suspensions were lower than  $\alpha$ , indicating attenuated protonation for kaolinite/montmorillonite and enhanced protonation for goethite/kaolinite and goethite/montmorillonite mineral systems. The higher the acidity of sites the more protons are exchanged for Hg(II). Therefore, mineral mixing reduced the acidity of reactive sites for kaolinite/montmorillonite mixed mineral phase. This may be due to the competition for the sorbing ion by the mixed mineral surfaces [3, 34]. The H<sup>+</sup>/-M exchange stoichiometry of <2 (Table 2) for Hg(II) sorption on both the single and mixed mineral phases agreed with previous findings [32]. This suggested that surface charges became increasingly changed as Hg(II) sorption progressed. In addition, mineral surfaces with few displaceable H<sup>+</sup> with limited pH-dependent sorption supported lower H<sup>+</sup>/-M stoichiometry.

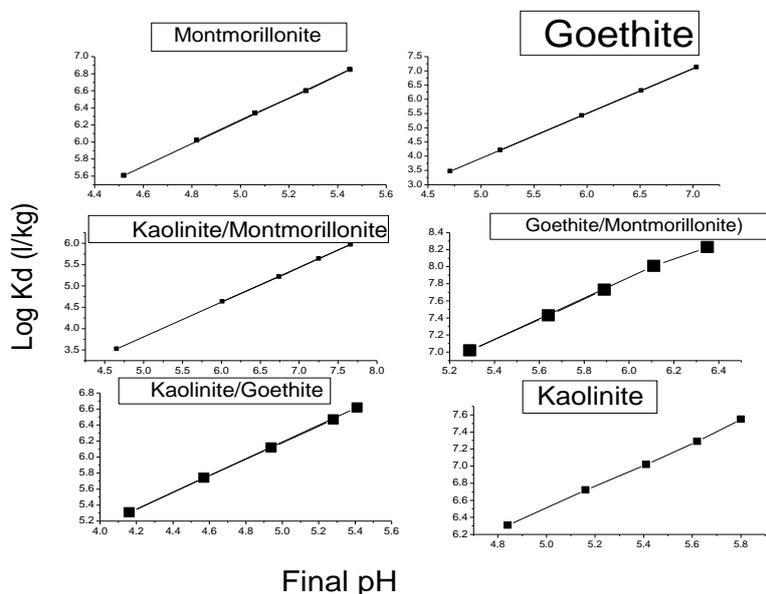


Figure 2: Plots of logkd versus Final pH for Hg(II) sorbed on single and mixed mineral systems.

### 3.2. Mixed mineral systems and sorption kinetics

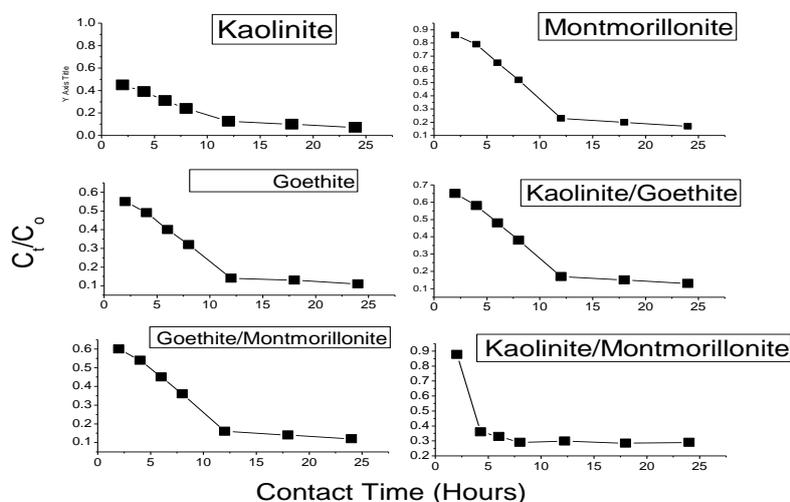
Sorption kinetics indicated a two-phase reaction probably attributed to outer sphere and inner sphere complexation with minimal intra-particle diffusion (Table 3, Fig 3). Mass transfer rate for the first-phase

reaction ( $K_{fI}$ ) decreased in the order goethite<kaolinite<montmorillonite for all single mineral systems. Mixed mineral systems had a decreasing mass transfer rate for montmorillonite-goethite<kaolinite-montmorillonite<kaolinite-goethite. Theoretical mass transfer rates were higher for kaolinite-montmorillonite and montmorillonite-goethite mixed mineral systems. Kaolinite-goethite exhibited a near similar theoretical  $K_{fI}$  value when compared with the actual  $K_{fI}$  value. This means that mineral mixing based on our empirical model decreased the mass transfer rates for kaolinite-montmorillonite and montmorillonite-goethite, not affecting kaolinite-goethite. For  $K_{fII}$  depicting the second-phase reaction, mass transfer rates decreased in the order, goethite<kaolinite<montmorillonite for the single mineral phases. For the mixed mineral systems, the order was, montmorillonite-goethite<kaolinite-goethite<kaolinite-montmorillonite. Based on the empirical model, mineral mixing did not change the mass transfer rates for the mixed mineral systems except for montmorillonite-goethite. There was a decrease in the mass transfer rate for montmorillonite-goethite. This is because the theoretical mass transfer rate for this mixed mineral system was higher than the actual mass transfer rate. Differences in mass transfer rates for Hg(II) ions transferred to the mineral reactive sites may be attributed (a) to different types of reactive sites on the single and mixed mineral systems (b) differences in BET surface area for the mineral systems [Fig 4] and (c) differences in particle size distribution of these mineral systems [ 6].

**Table 3: Mass transfer rates for Hg(II) sorbed on clay and (hydr)oxide mineral suspensions**

Parameters/ Mineral	Kaolinite	Montmorillo nite	Goethite	Kaolinite/ Montmorillo nite	Kaolinite/ Goethite	Montmorillo nite/ Goethite
slopeI(hr <sup>-1</sup> )	-0.07	-0.14	- 0.088	-0.062	-0.104	-0.097
slopeII(hr <sup>-1</sup> )	- 0.006	-0.011	- 0.007	-0.058	-0.0083	-0.0075
Exposed Surface Area (cm <sup>-1</sup> )	4700	1000	7100	8800	7900	14700
$K_{fI}$ (cmhr <sup>-1</sup> )	1.49 * <sup>-5</sup>	0.00014	1.24 * <sup>-5</sup>	7.04 * <sup>-6</sup>	1.31 * <sup>-5</sup>	6.6 * <sup>-6</sup>
Theoretical $K_{fI}$ (cmhr <sup>-1</sup> )	⊗	⊗	⊗	0.00007745	0.00001365	0.000082
$K_{fII}$ (cmhr <sup>-1</sup> )	1.28 * <sup>-6</sup>	0.000011	9.85 * <sup>-7</sup>	6.6 * <sup>-6</sup>	1.05 * <sup>-6</sup>	5.1 * <sup>-7</sup>
Theoretical $K_{fII}$ (cmhr <sup>-1</sup> )	⊗	⊗	⊗	0.00000614	0.00000113	0.0000059925

Note: ⊗ Not applicable



**Fig. 3. Effects of Hg(II) sorption on mass transfer rates for single and mixed mineral systems**

### 3.3 Mixed mineral systems and Hg(II) removal

The mercury uptake decreased as BET surface area increased for all single and mixed mineral systems within the range of pH studied. Furthermore, the mercury uptake does not appear to be linearly correlated with BET surface area, implying that mercury sorption onto single and mixed mineral systems is not strictly a surface phenomenon [30]

Mercury sorption exhibited a sinusoidal pattern probably attributed to outer sphere, inner sphere complexation and minimal intra-particle diffusion (Fig 5). All mineral systems demonstrated increase in % sorption at the onset of reaction, indicating a reaction dip after 6 hours of contact or residence time for kaolinite and kaolinite-montmorillonite. Goethite, goethite-montmorillonite and kaolinite-goethite exhibited a reaction dip at 12h residence time. There appeared to be no significant change in mercury sorption for these mineral systems after 12h contact time. Montmorillonite exhibited an astronomical Hg removal at the onset of reaction, experiencing a reaction dip after 8h contact time, then progressively decreased for the rest of the contact time. Differences in mercury removal as exhibited by these mineral systems could be attributed to differences in BET surface area (Table 1, Fig 4). Differences between actual and theoretical % sorption progressed from negative to positive for all mixed minerals as contact time was increased. This means that increase in contact time during Hg-sorbent interaction could help increase mercury removal based on mineral mixing (Fig 6).

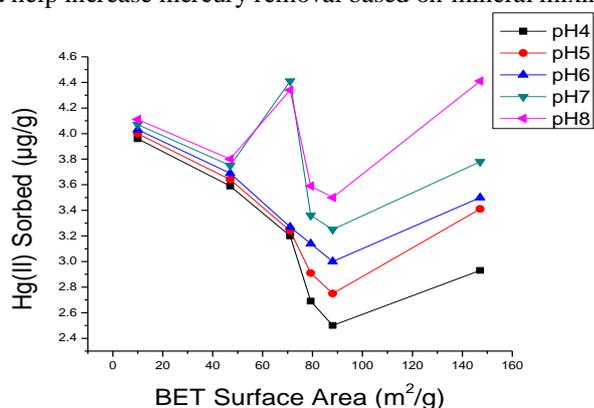


Figure 4: Plot of Hg(II) Sorbed vs. BET Surface area of clays and hydroxides [Table 1] at variable pH

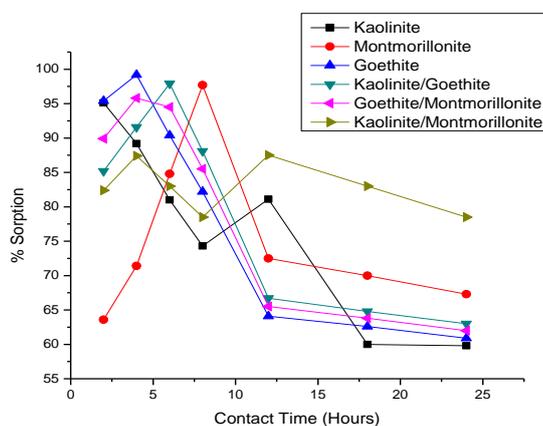
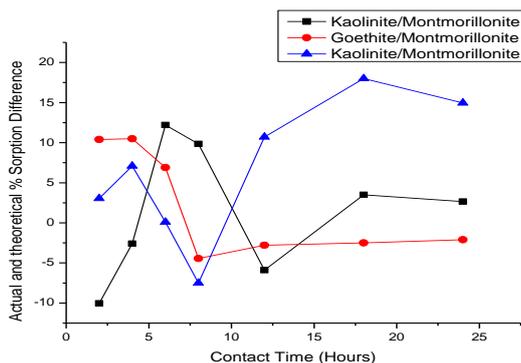


Fig.5. Plots of Hg(II) % sorption vs. time (hours) for single and mixed mineral systems.



**Fig. 6: Plots of differences between actual percent of Hg (II) removal and Theoretical percent vs. contact time (hours) for mixed mineral systems.**

#### IV. Conclusions

The reactivity and removal kinetics of Hg(II) treated with single and mixed mineral systems of kaolinite, montmorillonite and goethite have been investigated. Using empirical models derived from Freundlich isotherm model, proton coefficient for mercury sorbed on All proton coefficient for Hg(II) sorbed on single mineral systems except for goethite are greater than one. This indicated high level of protonation during the sorption process. Except for kaolinite/montmorillonite, mixed mineral suspension,  $\alpha_{total}$  for Hg(II) sorbed on the mixed mineral suspensions were lower than  $\alpha$ , indicating attenuated potonation for kaolinite/montmorillonite and enhanced protonation for goethite/kaolinite and goethite/montmorillonite mineral systems.

Mass transfer rate for the first-phase reaction ( $K_{f1}$ ) decreased in the order goethite<kaolinite<montmorillonite for all single mineral systems. Mixed mineral systems had a decreasing mass transfer rate for montmorillonite-goethite<kaolinit-montmorillonite<kaolinite-goethite. This suggested that different reactive sites were involved in the removal kinetics of the sorbing ions as sorption progressed. Differences in sorption kinetics between the single and mixed mineral phases may be attributed to differences in the BET surface area of single and mixed mineral systems. All mineral systems demonstrated increase in % sorption at the onset of reaction, indicating a reaction dip after 6 hours of contact or residence time for kaolinite and kaolinite-montmorillonite. Goethite, goethite-montmorillonite and kaolinite-goethite exhibited a reaction dip at 12h residence time. Differences between actual and theoretical % sorption progressed from negative to positive for all mixed minerals as contact time was increased. This means that increase in contact time during Hg-sorbent interaction could help increase mercury removal based on mineral mixing.

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